

Application of 2D-GCMS reveals many industrial chemicals in airborne particulate matter

Alam, M.S.; West, C.E.; Scarlett, A.G.; Rowland, S.J.; Harrison, R.M.

DOI:

[10.1016/j.atmosenv.2012.10.014](https://doi.org/10.1016/j.atmosenv.2012.10.014)

License:

Other (please specify with Rights Statement)

Document Version

Peer reviewed version

Citation for published version (Harvard):

Alam, MS, West, CE, Scarlett, AG, Rowland, SJ & Harrison, RM 2013, 'Application of 2D-GCMS reveals many industrial chemicals in airborne particulate matter', *Atmospheric Environment*, vol. 65, pp. 101-111.
<https://doi.org/10.1016/j.atmosenv.2012.10.014>

[Link to publication on Research at Birmingham portal](#)

Publisher Rights Statement:

NOTICE: this is the author's version of a work that was accepted for publication in *Atmospheric Environment*. Changes resulting from the publishing process, such as peer review, editing, corrections, structural formatting, and other quality control mechanisms may not be reflected in this document. Changes may have been made to this work since it was submitted for publication. A definitive version was subsequently published in *Atmospheric Environment*, [vol. 65, February 2013] DOI: <http://dx.doi.org/10.1016/j.atmosenv.2012.10.014>

Eligibility for repository : checked 6/08/2014

General rights

Unless a licence is specified above, all rights (including copyright and moral rights) in this document are retained by the authors and/or the copyright holders. The express permission of the copyright holder must be obtained for any use of this material other than for purposes permitted by law.

- Users may freely distribute the URL that is used to identify this publication.
- Users may download and/or print one copy of the publication from the University of Birmingham research portal for the purpose of private study or non-commercial research.
- User may use extracts from the document in line with the concept of 'fair dealing' under the Copyright, Designs and Patents Act 1988 (?)
- Users may not further distribute the material nor use it for the purposes of commercial gain.

Where a licence is displayed above, please note the terms and conditions of the licence govern your use of this document.

When citing, please reference the published version.

Take down policy

While the University of Birmingham exercises care and attention in making items available there are rare occasions when an item has been uploaded in error or has been deemed to be commercially or otherwise sensitive.

If you believe that this is the case for this document, please contact UBIRA@lists.bham.ac.uk providing details and we will remove access to the work immediately and investigate.

Application of 2D-GCMS Reveals Many Industrial Chemicals in Airborne Particulate Matter

Mohammed S. Alam,¹ Charles E. West²,
Alan G. Scarlett², Steven J. Rowland² and
Roy M. Harrison^{1*†}

¹Division of Environmental Health & Risk Management
School of Geography, Earth & Environmental Sciences
University of Birmingham, Edgbaston
Birmingham, B15 2TT, United Kingdom

²School of Geography, Earth and Environmental Sciences
University of Plymouth, Plymouth, Devon PL4 8AA
United Kingdom

* To whom correspondence should be addressed (Tel: +44 121 414 3494; Fax: +44 121 414 3709;
Email: r.m.harrison@bham.ac.uk)

†Also at: Department of Environmental Sciences / Center of Excellence in Environmental Studies, King Abdulaziz University, PO Box 80203, Jeddah, 21589, Saudi Arabia

1 **ABSTRACT**

2 Samples of airborne particulate matter (PM_{2.5}) have been collected in Birmingham, UK and
3 extracted with dichloromethane prior to analysis by two-dimensional GC separation and TOFMS
4 analysis. Identification of compounds using the NIST spectral library has revealed a remarkable
5 diversity of compounds, some of which have not been previously reported in airborne analyses.
6 Groups of compounds identified in this study include a large number of oxygenated VOC including
7 linear and branched compounds, substituted aromatic compounds and alicyclic compounds,
8 oxygenated polycyclic aromatic and alicyclic compounds, organic nitrogen compounds, branched
9 chain VOC and substituted aromatic VOC, phthalates, organo-phosphates and organo-sulphate
10 compounds. Many of the compounds identified are mass production chemicals, which due to their
11 semi-volatility enter the atmosphere and subsequently partition onto pre-existing aerosol. Their
12 contribution to the toxicity of airborne particulate matter is currently unknown but might be
13 significant. The diverse industrial uses and potential sources of the identified compounds are
14 reported.

15

16 Keywords: 2D GCMS, airborne chemicals, industrial chemicals, plasticisers, fire retardants

17

18

INTRODUCTION

Atmospheric aerosol particles are complex mixtures of nonvolatile and semi-volatile inorganic and organic compounds, directly released from a variety of sources, including vehicle emissions, combustion and vegetation, and may also be formed as products of gas to particle conversion through atmospheric reactions. These atmospheric aerosols affect the global radiation balance both directly and indirectly, by absorbing and scattering solar radiation, and by acting as cloud condensation nuclei, respectively. The determination of aerosol particle composition is therefore of significant interest owing to the potential impacts upon climate and human health (Zhang et al., 2012; Harrison and Yin, 2000).

Since a substantial fraction of atmospheric aerosol consists of thousands of organic compounds with a broad range of solubilities, reactivities and physical properties, identification of these species remains a challenge, with many compounds remaining unidentified (Goldstein et al., 2007). Added complexities to the analyses of these aerosols are the trace-level concentrations of organics and difficulties in sampling. The majority of studies have collected aerosol samples onto filters, where the compounds of interest are removed by solvent extraction or thermal desorption, followed by analysis with gas chromatography coupled to mass spectrometry (GC-MS) (Tomashuk et al., 2012; Yin et al., 2010; Harrison and Yin, 2010). However, owing to the large variety of compounds present in samples, conventional GC often proves inadequate for the separation of components, thus leading to severe peak co-elutions and poor chromatography resulting from insufficient resolving power.

Some studies have made extensive use of one-dimensional GC-MS together with selective extractions and derivatisation procedures to identify large numbers of compounds in airborne particulate matter. These include Pio et al. (2001) who identified around 400 individual organic compounds in particulate matter from a Portuguese rural area. By specifically targeting

homologous series and functional groups, they were able to identify aliphatics, polycyclic aromatic hydrocarbons (PAHs), oxy-PAH, ketones, aldehydes, alcohols, n-alkanoic acids, n-alkenoic acids, aliphatic dicarboxylic acids, terpenic constituents and some VOC oxidation products. In a study of particulate matter sampled in Kuala Lumpur, Malaysia, Omar et al. (2007) reported concentrations of PAH (14 compounds), petroleum molecular markers (4), n-alkanols (7), triterpenols (6), n-alkanoic acids (7), soil and biomass burning carbohydrates (1), sterols (5), products of plastics combustion (3), n-alkanoic acids (4) and mono and diacylglycerols (3). In all, within samples from the urban area, the specific compounds analysed accounted for $5.3 \mu\text{g m}^{-3}$ out of a total solvent extractable organic matter concentration of $18.1 \mu\text{g m}^{-3}$.

Recent advances in comprehensive two-dimensional gas chromatography (GC \times GC) have led to its successful application in measuring a wide range of volatile and semi-volatile organic compounds in air (Adahchour et al., 2008; Laitinen et al., 2010; Arsene et al., 2011). For example, the unresolved complex mixture (UCM) clear from the white line in Figure 1(A) in a traditional 1D GCMS separation typically makes up about 95% of the area of the chromatogram and is usually unresolved by analysis. 2D GCMS techniques provide an enormous advance in terms of resolving individual peaks but have yet to be exploited very fully in atmospheric analysis. Arsene et al. (2011) have recently reviewed the application of GC \times GC-MS hyphenated techniques for the analysis of volatile and semi-volatile compounds in air. They report only 12 published studies of atmospheric measurements, which are focussed largely on lower molecular mass species (less than C_{12}), oxygenates, or forest atmospheres. The studies demonstrate applicability to alkanes up to C_{30} (Laitinen et al., 2010) and roadside nanoparticles (Ochiai et al., 2007).

GC \times GC was first exploited in atmospheric analysis by Lewis et al. (2000), who isolated and classified more than 500 VOCs in urban air. This powerful technique has since been combined with flame ionisation detection (FID) and with time of flight mass spectrometry (TOFMS) in several

1 studies (Pratt et al., 2012). For example, Xu and co workers (Xu et al., 2003a,b) utilised GC×GC–
2 TOFMS to measure C₇–C₁₁ aromatic and *n*-alkane hydrocarbons in air, isolating approximately 650
3 individual peaks of which 235 were identified using published retention indices. Lee et al. (2006)
4 exploited the use of GC×GC–FID to study VOC ratios of high molecular mass aromatics which
5 could not be speciated using conventional GC–MS. Welthagen et al. (2003) and Hamilton et al.
6 (2004) used direct thermal desorption GC×GC–TOFMS for the characterisation of semi-volatile
7 organic compounds in PM_{2.5} detecting (but not identifying) 15,000 and 10,000 peaks, respectively.
8 Of these 10,000 peaks, Hamilton et al. (2004) report the positive identification of 130 OVOCs,
9 where a further 100 oxygenated species were observed but not positively identified owing to the
10 lack of mass spectral library or pure components. Williams et al. (2010) reported the first hourly in-
11 situ measurements of speciated organic aerosol composition in an urban environment, where
12 approximately 300 nonvolatile and semivolatile organic compounds were identified using thermal
13 desorption. More recently, this technique was further improved by coupling GC×GC to thermal
14 desorption aerosol gas chromatography and has been deployed in both chamber studies
15 investigating the photo-oxidation of diesel exhaust (Isaacman et al., 2011) and field studies (Worton
16 et al., 2012), allowing real time measurements of a diverse range of compound classes.

17
18 2D GC-MS has extended knowledge of organic compounds present in the atmosphere particularly
19 in the directions of organo-nitrogen compounds and oxidised compounds. However, the knowledge
20 that thousands of VOC and SVOC compounds exist in samples that are collected from various
21 studies and are yet to be identified and routinely measured, provides motivation for this study.
22 GC×GC-TOFMS was applied to identify organic compounds in atmospheric aerosol collected in
23 Birmingham, UK. In this study, we have not reported groups of compounds such as *n*-alkanes and
24 PAH which have been widely reported before, but have focussed particularly on compound types
25 which have rarely, if ever, been previously reported in the atmosphere by selecting peaks for

characterisation which were not part of the main homologous series identifiable from the chromatograms.

EXPERIMENTAL

Site Location

Samples of airborne fine particulate matter (PM_{2.5}) were collected at the Elms Road Observatory Site (EROS, 1.93°W; 52.45°N). EROS is an urban background site within the University of Birmingham campus. The site is approximately 3.5 km southwest of Birmingham city centre and is subject to anthropogenic emissions from a nearby railway, some moderately busy roads and surrounding activities from local residents.

Sampling

Daily PM_{2.5} samples were collected for 5 days (9-14 February 2011) using a high volume air sampler TE-6070 (Tisch Environmental, Inc). The high volume sampler involves drawing large volumes of air (800-1000 m³) through a 20 × 25 cm quartz fibre filter (QMA) substrate. Filters were pre-heated at 550°C for 12 hours and stored in a freezer prior to sampling. After exposure the filters were enclosed with clean pre-heated foil and stored in an air tight metal box under conditions of approximately -18°C prior to analysis. No attempt was made to evaluate or correct for sampling artefacts, which can be both positive (adsorption of vapour to the filter or collected particles) or negative (loss by volatilisation).

Extraction and Analysis

Whole quartz fibre filters were extracted with two 30 ml aliquots of DCM (HPLC grade) and methanol (HPLC grade) sequentially. The extracts were combined, filtered and concentrated using a turbo evaporator and then blown down under a gentle nitrogen flow to a final volume of 250 µL. Comprehensive two-dimensional gas chromatography/time of flight mass spectrometry (GC×GC-

TOFMS) was utilised to analyse the sample extracts with conditions similar to that reported by Rowland et al. (2012). Briefly, analyses were conducted using an Agilent 7890A gas chromatograph (Agilent Technologies, Wilmington, DE, USA) equipped with a Zoex ZX2GC×GC cryogenic modulator (Houston, TX, USA) interfaced with a Almsco BenchToFdx™ time of flight mass spectrometer (Almsco International, Lantrisant, UK) operated in positive ion electron ionisation mode and calibrated with perfluorotributylamine. The scan speed was 50 Hz. The first-dimension column was 50 m × 0.25 mm × 0.40 µm VF1-MS (Varian, Palo Alto, CA, USA). The second-dimension column was 2.5 m × 0.15 mm × 0.15 µm VF-17MS (Varian). 1 µL of sample was injected in a splitless mode at 280°C. The initial temperature of the oven (60°C) was held for 1 minute and then increased at 5°C min⁻¹ to 325°C and held for 10 minutes. The modulation period was 4 seconds. The transfer line temperature was 280°C and the ion source temperature was 300°C. Helium was used as the carrier gas at a constant flow rate of 0.8 mL min⁻¹. Subsequent data processing was conducted using GC Image™ v2.1 (Zoex).

RESULTS AND DISCUSSION

Five high volume urban aerosol samples collected in Birmingham, on 9–14 February 2011 were analysed utilising GC×GC-TOFMS, to identify compounds within the collected samples. A typical two-dimensional separation and subsequent total ion count is presented in Figure 1(B), where each coloured spot represents an individual compound with a corresponding full mass spectrum. The intense yellow spots are major peaks, while the magenta and dark blue spots are medium and small peaks, respectively. Compounds were separated on a volatility basis in the first dimension (X-axis) and polarity in the second (Y-axis). Approximately 1,500 peaks were found in the two-dimensional chromatogram, using the blob detection routine performed by GC Image v2.1b3 (Zoex Corporation, Houston, US) software.

Peak identification was based on mass spectral data from the NIST library, where two match factors were used to assess the similarity between the library and measured mass spectra. These similarity factors were: (i) match factor (MF) – described how well the library equivalent matched the entire mass spectra of the selected chromatography peak, and (ii) reverse match factor (RMF) – described how well the explicit masses present in the library equivalent matched the masses measured for the selected peak. Positive identification was based on MF and RMF of above 750 and 800, respectively, consistent with previous studies (Dalluge et al., 2002; Ozel et al., 2009).

Although a large number of peaks were unidentifiable owing to insufficient data available in the mass spectral library, various compound groups including alkanes, alkenes, cycloalkanes, aldehydes, ketones, carboxylic acids and esters, PAHs, oxygenated-PAHs and substituted aromatics were positively identified. The ordered appearance of chemically related compounds in the GC×GC chromatogram enables further identification of species owing to the physicochemical similarities within compound classes and their steady changes with increasing chain length and/or molecular sizes. This allows the identification of unknown species without unique mass spectra, as demonstrated for the dihydro-furanone series by Hamilton et al. (2004) and amides by Ozel et al. (2009).

Further information regarding the identification of compound classes can be achieved by producing GC×GC contour plots for selected mass fragments. An example of a selected ion chromatogram is shown in Figure 1(C), for m/z 85 (a branched alkane fragment), demonstrating the increased resolution and ordered appearance of chemically related compounds. Mass pattern selection rules were developed by Welthagen et al. (2003) in order to group peaks into distinct chemical classes, facilitating interpretation of the peaks detected in PM_{2.5} samples. These selection rules provided an estimate for the identification of specific compound classes that have been extensively reported in

earlier studies (see Welthagen et al., (2003) for ion abundance rules). Their retention time selection rules, however, were not absolute and therefore were not taken into consideration in this study.

Figures 2, 3 and 4 illustrate the positive identification of a range of compounds from both natural and industrial sources, based on the NIST spectral library match. The measured mass spectra demonstrate the excellent agreement between our spectra and the reference spectra contained within the NIST library. Example spectra for some naturally occurring compounds are illustrated in Figure 2; these include three methyl esters emitted from pine resin and plants (Figure 2, A-F) and *cis*-pinonic acid (Figure 2, G-H), an atmospheric organic compound produced by the oxidation of terpenes (Li et al. 2010). The spectra contained within Figure 3, demonstrate the positive identification of industrial chemicals used in PVC (Figure 3, A-B), other plastics (Figure 3, C-F) and nylon (Figure 3, G-H). Caprolactam (shown in Figure 3, G-H) was present in four out of five samples. Other industrial compounds identified are shown in Figure 4, which include chemicals used in paper production (Figure 4, A-B), as non ionic surfactants (Figure 4, C-D), as pesticides (Figure 4, E-F) and as flame retardants (Figure 4, G-H). All positively identified compounds were further confirmed by assessing the consistency between the two retention times measured within the five different samples.

In the ambient PM_{2.5} samples collected at Birmingham, UK, 100 out of the 1,500 peaks detected were qualitatively identified, exclusive of *n*-alkanes (C₁₀ – C₃₁), *n*-alkanals (C₁₈ – C₂₉), *n*-alkanoic acids (C₅ – C₂₅), furanones and polycyclic aromatic hydrocarbons (PAHs), all of which have been extensively reported in earlier studies. The compounds reported in this study were present in at least three of the five samples collected and have been split into five groups: (i) Oxygenated volatile organic compounds (OVOC), Table 1; (ii) Organic nitrogen compounds (ON), Table 2; (iii) Oxygenated polycyclic aromatic and alicyclic hydrocarbons, Table 3; (iv) other semi-volatile organic compounds (VOC), Table 4; and (v) Phthalates, organic phosphates and organic sulphates,

1 Table 5. All compounds are reported with their corresponding retention times, MF, RMF and
2 industrial use. The postulated industrial uses of the compounds listed in Tables 1-5 are not exclusive
3 primary sources for each compound, as many of these compounds may also be products of
4 secondary reactions in the atmosphere. For example 9,10-anthracenedione may be a primary
5 component of vehicle emissions (Jakober et al., 2007) and is utilised in the production of dyes,
6 paper and pesticides; but is also a product from the oxidation reaction of its parent PAH compound,
7 anthracene, in the atmosphere. The industrial uses and potential sources of the identified
8 compounds were gathered using internet search engines, emission databases and Wikipedia.
9 Compounds for which little information was available in the literature were traced back to
10 manufacturers to deduce a potential industrial uses / sources.

11
12 As explained for 9,10-anthracenedione, a component of organic aerosol (OA) can have a direct
13 source of emission and is then termed as primary organic aerosol (POA), or be formed via
14 partitioning of oxidised gas phase species, when it is known as secondary organic aerosol (SOA).
15 OA compounds therefore have a wide range of organic functionalities and volatilities, making it
16 extremely difficult to distinguish between primary and secondary sources in PM samples, without
17 the aid of source apportionment techniques such as chemical mass balance (Yin et al., 2010),
18 principal component analysis (Mari et al., 2010) and/or positive matrix factorisation (Williams et
19 al., 2010). Hamilton et al. (2004) identified 130 OVOC in urban aerosol using direct thermal
20 desorption (DTD) coupled to GC×GC-TOFMS, based on retention behaviour and spectral matches.
21 They collected their samples at both roadside locations in London and on a rooftop close to Leeds
22 city centre. Many of the carbonyl species observed in their study could be mechanistically linked to
23 gas phase aromatic hydrocarbon oxidation, and hence could be secondary in nature. Many of the
24 compounds detected by Hamilton et al. (2004) were not identified in the aerosol samples collected
25 in this study, which may be due to differing sampling locations and local atmospheric composition.
26 Roadside sampling sites include fresh gasoline and diesel vehicle emissions of VOCs, many of

1 which may be oxidised rapidly, dependent upon gas-phase and heterogeneous reaction rates with
2 respect to atmospheric oxidants (George et al., 2010). The majority of OVOCs observed in this
3 study are widely used in industry (see Table 1), but may also be secondary in origin owing to long
4 range transport and atmospheric processing. Of increasing interest is 7,9-di-tert-butyl-1-
5 oxaspiro(4,5)deca-6,9-diene-2,8-dione (see Figure 3, E-F), a degradation product of antioxidants
6 used as additives for plastic materials and engine oils, which has recently been identified as a water
7 contaminant due to leaching from water pipes. To date, however, its toxicity has not been assessed
8 and may be of significance.

9
10 Twelve of the 16 priority PAHs, defined by the United States Environment Protection Agency
11 (EPA), were observed in more than three of the samples. Four PAHs that were not identified in any
12 of the samples were anthracene, indeno(1,2,3-cd)pyrene, dibenz(a,h)anthracene and coronene.
13 Although the absence of these PAHs may be due to them having lower emission rates relative to
14 other PAHs, their fast conversion to PAH derivatives by atmospheric oxidants cannot be ruled out.
15 The gas phase reaction of anthracene + OH is reported as approximately $1 \times 10^{-10} \text{ cm}^3 \text{ molecules}^{-1} \text{ s}^{-1}$
16 (Goulay et al., 2005), which is found to be an order of magnitude slower than the more abundant
17 PAH, phenanthrene. This relatively fast conversion of anthracene to its corresponding quinone
18 compound may therefore partially explain the absence of anthracene and explain the presence of
19 9,10-anthracenedione in all samples, as well as primary emissions from vehicles and the production
20 of dyes, paper and pesticides in the area. Recently, indeno(1,2,3-cd)pyrene, which is known as a
21 relatively stable PAH, has been shown to undergo considerable degradation when exposed to
22 atmospheric oxidants and was also not observed in any of the samples collected in this study
23 (Ringuet et al., 2012). The compounds, [1R-(1 α ,4 α β ,10 $\alpha\alpha$)]-1,2,3,4,4a,9,10,10a-octahydro-1,4a-
24 dimethyl-7-(1-methylethyl)-1-phenanthrenecarboxylic acid, methyl ester (see Figure 2, A-B); 7-
25 oxodehydroabietic acid (dehydroabietic acid), methyl ester (see Figure 2, C-D) and 16 β H-kauran-
26 16-ol, however, are found naturally in the environment (see Table 2).

1 While there have been extensive studies of oxygenated organics in the atmosphere, relatively little
2 attention has been given to organic nitrates, despite their environmental importance. A number of
3 ON species have been listed by the EPA as hazardous air pollutants (*e.g.* nitro-PAHs), yet the lack
4 of satisfactory methods for analysing ON compounds means that there is limited information in the
5 literature. Although no nitro-PAHs were identified in this study, possibly owing to their low
6 concentration in the samples, only a few ON compounds were identified in three or more samples
7 (Table 3). Hexadecanamide was present in all samples, in agreement with the results of Ozel et al.
8 (2009), who utilised direct thermal desorption (DTD) in combination with GC×GC-TOFMS to
9 analyse ON compounds in urban aerosol samples from the same site. DTD permits the analysis of
10 samples without any prior solvent extraction or sample preparation, where the samples are placed
11 directly onto a desorption tube and subjected to controlled heating. The desorbed volatiles are then
12 injected directly into the column. The authors reported between 17 and 57 ON compounds in any
13 one sample, with the lowest being recorded during February. The modest number of ON
14 compounds reported in Table 3 suggests that the chemical composition of the samples collected
15 here varies from day to day, owing to changing meteorology and complexity of emission sources,
16 consistent with Ozel et al. (2009). A further reason for the discrepancy between the number of ON
17 compounds reported by Ozel et al. (2009) and those reported in this study may be due to the
18 differing analytical procedures employed.

19

20 Table 4 reports the detection of a number of hydrocarbons, not commonly reported as present in the
21 atmosphere. For example, 2,6,10,14-tetramethyl-hexadecane, (or phytane) has been used as a
22 biomarker in petroleum geochemistry, where ratios of pristane and phytane have been exploited as
23 indicators of maturity and extent of biodegradation (Hunt et al., 2002), but to our knowledge has not
24 been routinely measured in atmospheric / environmental studies. Hexatriacontane occurs in
25 lubricating oils manufactured within the petroleum industry; thus the presence of hexatriacontane in

all five aerosol samples indicates that routine measurements may benefit future studies, serving as an oil tracer determining the unburned oil contribution to PM (Zielinska et al., 2008).

Table 5 shows some of the many compounds which are used in industry, as plasticizers, flame retardants and paint solvents. These compounds have received less attention in airborne analyses. Phthalates and organophosphates are ubiquitous chemicals in the environment and are used commercially as plasticizers and flame retardants, but are not usually measured in the atmosphere. The majority of these compounds were detected in all the samples, with abundances considerably larger than that found in the filter blank (compounds found in the filter blank are shown with an asterisk in Table 5). Some studies have evaluated the concentration of phthalates and organophosphates in indoor air (Bergh et al., 2011), but relatively few studies have focused on levels in outdoor environments owing to suspected levels being an order of magnitude lower (Rudel et al., 2009). Although volatilisation from environmental surfaces such as soil may not be a significant source to the atmosphere (Harrad et al., 2006), the adsorption of these compounds by dust particles may be responsible for the widespread distribution of these chemicals. Furthermore, dependant upon volatility, phthalates, organophosphates and organosulphates may be adsorbed to airborne particulate matter and may be subject to degradation by atmospheric oxidants, such as O_3 , OH and NO_3 . The rate of degradation of these compounds, both in the vapour and condensed phases, by atmospheric oxidants has not been experimentally measured. However, half lives of 1-100 days have been estimated by using structure activity relationship methods for phthalates in the vapour phase with respect to OH (Atkinson et al., 1987; Meylan et al., 1992). These predicted half lives suggest that oxidation by OH radicals could be the major degradation pathway, transforming these plasticizers released into the atmosphere and could subsequently form further products of varying toxicity.

1 **CONCLUSIONS**

2 Earlier GC-MS studies of semi-volatile compounds in atmosphere aerosol have focused on specific
3 groups of compounds and have therefore not captured the remarkable range of compounds present.
4 This study has deliberately taken the opposite line of ignoring those groups of compounds for which
5 there are a number of reports in the literature, but focussing on unrelated compounds in order to get
6 a measure of the diversity of those present. As a consequence, it has revealed a quite exceptionally
7 diverse range of compounds, many of which are in use by industry and incorporated into consumer
8 products, for example, as plasticisers or fire retardants. Given the vapour pressures of such
9 substances, it is not surprising that small amounts enter the atmosphere and future research on
10 airborne particulate matter and health needs to take account of their presence. While individually
11 they are unlikely to represent a toxic hazard, collectively, they may contribute significantly to the
12 toxicity of airborne particulate matter. It seems far more likely that the phthalate plasticisers, for
13 which there is evidence of endocrine disruptive potential (Sharpe and Irvine, 2004), or organo-
14 phosphate fire retardants, present a significantly greater contribution to toxicity than the alkanes
15 which are more commonly measured. While the polycyclic aromatic hydrocarbons probably
16 contribute the majority of the carcinogenicity of urban particulate matter, contributions from some
17 of the other compounds present cannot be ruled out.

18

19 **ACKNOWLEDGEMENT**

20 Funding for part of this research was provided by an Advanced Investigators Grant (no. 228149)
21 awarded to SJR for project OUTREACH, by the European Research Council, to whom we are
22 extremely grateful. We also acknowledge funding from the Natural Environment Research Council
23 (to RMH and MSA) and the assistance of Pallavi Pant and Jianxin Yin in sample collection and
24 extraction.

25

26

1

2

REFERENCES

- Adahchour, M., Beens, J. and Brinkman, U. A. T., 2008. Recent developments in the application of comprehensive two-dimensional gas chromatography. *Journal of Chromatography A*, 1186, 67-108.
- Arsene, C., Vione, D., Grinberg, N. and Olariu, R. I., 2011. GC x GC-MS hyphenated techniques for the analysis of volatile organic compounds in air. *Journal of Liquid Chromatography & Related Technologies*, 34, 1077-1111.
- Atkinson, R., 1987. A structure-activity relationship for the estimation of rate constants for the gas-phase reactions of OH radicals with organic compounds. *International Journal of Chemical Kinetics*, 19, 799-828
- Bergh, C., Torgrip, R., Emenius, G. and Ostman, C., 2011. Organophosphate and phthalate esters in air and settled dust – a multi-location indoor study. *Indoor Air*, 21, 67-76
- Dalluge, J., van Rijn, M., Beens, J., Vreuls, R. J. J. and Brinkman, U. A. T., 2002. Comprehensive two-dimensional gas chromatography with time-of-flight mass spectrometric detection applied to the determination of pesticides in food extracts. *Journal of Chromatography A*, 965, 207-217.
- George, I. J. and Abbatt, J. P. D., 2010. Heterogeneous oxidation of atmospheric aerosol particles by gas-phase radicals. *Nature Chemistry*, 2, 713-722.
- Goldstein, A. H. and Galbally, I. E., 2007. Known and unexplored organic constituents in the earth's atmosphere. *Environmental Science & Technology*, 41, 1514-1521.
- Goulay, F., Rebrion-Rowe, C., Le Garrec, J. L., Le Picard, S. D., Canosa, A. and Rowe, B. R., 2005. The reaction of anthracene with OH radicals: An experimental study of the kinetics between 58 and 470 K. *Journal of Chemical Physics*, 122 , 10, 104308
- Hamilton, J. F., Webb, P. J., Lewis, A. C., Hopkins, J. R., Smith, S. and Davy, P., 2004. Partially oxidised organic components in urban aerosol using GCXGC-TOF/MS. *Atmospheric Chemistry and Physics*, 4, 1279-1290.
- Harrad, S. and Hunter, S., 2006. Concentrations of polybrominated diphenyl ethers in air and soil on a rural-urban transect across a major UK conurbation. *Environmental Science & Technology*, 40, 4548-4553.
- Harrison, R. M. and Yin, J. X., 2010. Chemical speciation of PM_{2.5} particles at urban background and rural sites in the UK atmosphere. *Journal of Environmental Monitoring*, 12, 1404-1414.
- Harrison, R. M. and Yin, J., 2000. Particulate matter in the atmosphere: Which particle properties are important for its effects on health? *Science of the Total Environment*, 249, 85-101.
- Hunt, J. M., Philp, R. P. and Kvenvolden, K. A., 2002. Early developments in petroleum geochemistry. *Organic Geochemistry*, 33, 1025-1052
- Isaacman, G., Worton, D. R., Kreisberg, N. M., Hennigan, C. J., Teng, A. P., Hering, S. V., Robinson, A. L., Donahue, N. M. and Goldstein, A. H., 2011. Understanding evolution of product composition and volatility distribution through in-situ GCxGC analysis: a case study of longifolene ozonolysis. *Atmospheric Chemistry and Physics*, 11, 5335-5346

- 1 Jakober, C. A., Riddle, S. G., Robert, M. A., Destailats, H., Charles, M. J., Green, P. G. and
2 Kleeman, M. J., 2007. Quinone emissions from gasoline and diesel motor vehicles. *Environmental
3 Science & Technology*, 41, 4548-4554.
4
- 5 Laitinen, T., Martin, S. H., Parshintsev, J., Hyotylainen, T., Hartonen, K., Riekkola, M. L.,
6 Kulmala, M. and Pavon, J. P. P., 2010. Determination of organic compounds from wood
7 combustion aerosol nanoparticles by different gas chromatographic systems and by aerosols mass
8 spectrometry. *Journal of Chromatography*, 1217, 151-159.
9
- 10 Lee, J. D., Lewis, A. C., Monks, P. S., Jacob, M., Hamilton, J. F., Hopkins, J. R., Watson, N. M.,
11 Saxton, J. E., Ennis, C., Carpenter, L. J., Carslaw, N., Fleming, Z., Bandy, B. J., Oram, D. E.,
12 Penkett, S. A., Slemr, J., Norton, E., Rickard, A. R., Whalley, L. K., Heard, D. E., Bloss, W. J.,
13 Gravestock, T., Smith, S. C., Stanton, J., Pilling, M. J., Jenkin, M. E., 2006. Ozone photochemistry
14 and elevated isoprene during the UK heatwave of August 2003. *Atmospheric Environment*, 40,
15 7598-7613.
16
- 17 Lewis, A. C., Carslaw, N., Marriott, P. J., Kinghorn, R. M., Morrison, P., Lee, A. L., Bartle, K. D.
18 and Pilling, M. J., 2000. A larger pool of ozone-forming carbon compounds in urban atmospheres.
19 *Nature*, 405, 778-781.
20
- 21 Li, X., Hede, T., Tu, Y., Leck, C., Agren, H., 2010. Surface-active *cis*-pinonic acid in atmospheric
22 droplets: A molecular dynamics study. *Journal of Physical Chemistry Letters*, 4, 769-773
23
- 24 Mari, M., Harrison, R. M., Schuhmacher, M., Domingo, J. L. and Pongpiachan, S., 2010. Inferences
25 over the sources and processes affecting polycyclic aromatic hydrocarbons in the atmosphere
26 derived from measured data. *Science of the Total Environment*, 408, 2387-2393
27
- 28 Meylan, W. and Howard, P. H., 1992. Molecular topology/fragment contribution method for
29 predicting soil sorption coefficients. *Environmental Science & Technology*, 26, 1560-1567
30
- 31 Ochiai, M., Ieda, T., Sasamoto, K., Fushimi, A., Hasegawa, S., Tanabe, K. and Kobayashi, S., 2007.
32 Comprehensive two-dimensional gas chromatography coupled to high-resolution time-of-flight
33 mass spectrometry and simultaneous nitrogen phosphorous and mass spectrometric detection for
34 characterization of nanoparticles in roadside atmosphere. *Journal of Chromatography*, 1150, 13-20.
35
- 36 Omar, N. Y. M. J., Abas, M. R. B., Rahman, N. A., Tahir, N. M., Rishdi, A. I. and Simoneit, R. T.,
37 2007. Levels and distributions of organic source tracers in air and roadside dust particles of Kuala
38 Lumpur, Malaysia. *Environmental Geology*, 52, 1485-1500.
39
- 40 Ozel, M. Z., Ward, M. W., Hamilton, J. F., Lewis, A. C., Raventos-Duran, M. T. and Harrison, R.
41 M., 2009. Analysis of Organic Nitrogen Compounds in Urban Aerosol Samples Using GCxGC-
42 TOF/MS. *Aerosol Science and Technology*, 44, 109-116.
43
- 44 Pio, C. A., Alves, C. A. and Duarte, A. C., 2001. Identification, abundance and origin of
45 atmospheric organic particulate matter in a Portuguese rural area. *Atmospheric Environment*, 35,
46 1365-1375.
47
- 48 Pratt, K. A. and Prather, K. A., 2012. Mass spectrometry of atmospheric aerosols - Recent
49 developments & applications. Part I: Off-line mass spectrometry techniques. *Mass Spectrometry
50 Reviews*, 31, 1-16
51
52

- 1 Ringuet, J., Albinet, A., Leoz-Garziandia, E., Budzinski, H. and Villenave, E., 2012. Reactivity of
2 polycyclic aromatic compounds (PAHs, NPAHs and OPAHs) adsorbed on natural aerosol particles
3 exposed to atmospheric oxidants. *Atmospheric Environment*, 61, 15-22.
- 4
- 5 Rowland, S. J., West, C. W., Scarlett, A. G., Ho, C. and Jones, D., 2012. Differentiation of two
6 industrial oil sands process-affected waters by two-dimensional gas chromatography/mass
7 spectrometry of diamondoid acid profiles. *Rapid Communications in Mass Spectrometry*, 26, 572-
8 576.
- 9
- 10 Rudel, R. A. and Perovich, L. J., 2009. Endocrine disrupting chemicals in indoor and outdoor air.
11 *Atmospheric Environment*, 43, 170-181.
- 12
- 13 Sharpe, R. M. and Irvine, D. S., 2004. How strong is the evidence of a link between environmental
14 chemicals and adverse effects on human reproductive health? *British Medical Journal*, 328, 447-
15 451.
- 16
- 17 Tomashuk, T. A., Truong, T. M., Mantha, M. and McGowin, A. E., 2012. Atmospheric polycyclic
18 aromatic hydrocarbon profiles and sources in pine needles and particulate matter in Dayton, Ohio,
19 USA. *Atmospheric Environment*, 51, 196-202.
- 20
- 21 Welthagen, W., Schnelle-Kreis, J. and Zimmermann, R., 2003. Search criteria and rules for
22 comprehensive two-dimensional gas chromatography-time-of-flight mass spectrometry analysis of
23 airborne particulate matter. *Journal of Chromatography A*, 1019, 233-249.
- 24
- 25 Williams, B. J., Goldstein, A. H., Kreisberg, N. M., Hering, S. V., Worsnop, D. R., Ulbrich, I. M.,
26 Docherty, K. S. and Jimenez, J. L., 2010. Major components of atmospheric organic aerosol in
27 southern California as determined by hourly measurements of source marker compounds.
28 *Atmospheric Chemistry and Physics*, 10, 11577-11603
- 29
- 30 Worton, D. R., Kreisberg, N. M., Isaacman, G., Teng, A. P., McNeish, C., Gorecki, T., Hering, S.
31 V. and Goldstein, A. H., 2012. Thermal desorption comprehensive two-dimensional gas
32 chromatography: An improved instrument for in-situ speciated measurements of organic aerosols.
33 *Aerosol Science and Technology*, 46, 380-393.
- 34
- 35 Xu, X., van Stee, L. L. P., Williams, J., Beens, J., Adahchour, M., Vreuls, R. J. J., Brinkman, U. A.
36 T. and Lelieveld, J., 2003a. Comprehensive two-dimensional gas chromatography (GC x GC)
37 measurements of volatile organic compounds in the atmosphere. *Atmospheric Chemistry and
38 Physics*, 3, 665-682.
- 39
- 40 Xu, X., Williams, J., Plass-Dulmer, C., Berresheim, H., Salisbury, G., Lange, L. and Lelieveld, J.,
41 2003b. GCxGC measurements of C-7-C-11 aromatic and n-alkane hydrocarbons on Crete, in air
42 from Eastern Europe during the MINOS campaign. *Atmospheric Chemistry and Physics*, 3, 1461-
43 1475.
- 44
- 45 Yin, J., Harrison, R. M., Chen, Q., Rutter, A. and Schauer, J. J., 2010. Source apportionment of
46 fine particles at urban background and rural sites in the UK atmosphere. *Atmospheric Environment*,
47 44, 841-851.
- 48
- 49 Zielinska, B., Campbell, D., Lawson, D. R., Ireson, R. G., Weaver, C. S., Hesterberg, T. W.,
50 Larson, T., Davey, M. and Liu, L. J. S., 2008. Detailed characterization and profiles of crankcase
51 and diesel particular matter exhaust emissions using speciated organics. *Environmental Science &
52 Technology*, 42, 5661-5666

1
2 Zhang, R. J., Ho, K. F. and Shen, Z. X., 2012. The role of aerosol in climate change, the
3 environment, and human health. *Atmospheric and Oceanic Science Letters*, 5, 156-161.
4

TABLE LEGENDS

Table 1. Oxygenated volatile organic compounds detected in samples collected in Birmingham, UK, between 9 Feb – 14 Feb 2011

RMF – reverse match factor; MF – match factor; tR1 – retention time 1 (minutes); tR2 – retention time 2 (minutes)

Table 2. Oxygenated polycyclic aromatic and alicyclic compounds detected in samples collected in Birmingham, UK, between 9 Feb – 14 Feb 2011

RMF – reverse match factor; MF – match factor; tR1 – retention time 1 (minutes); tR2 – retention time 2 (minutes)

Table 3. Organic nitrogen compounds detected in samples collected in Birmingham, UK, between 9 Feb – 14 Feb 2011

RMF – reverse match factor; MF – match factor; tR1 – retention time 1 (minutes); tR2 – retention time 2 (minutes)

Table 4. Other semi-volatile organic compounds detected in samples collected in Birmingham, UK, between 9 Feb – 14 Feb 2011

RMF – reverse match factor; MF – match factor; tR1 – retention time 1 (minutes); tR2 – retention time 2 (minutes)

Table 5. Phthalates, organophosphates and organosulphate compounds detected in samples collected in Birmingham, UK, between 9 Feb – 14 Feb 2011

RMF – reverse match factor; MF – match factor; tR1 – retention time 1 (minutes); tR2 – retention time 2 (minutes)

* Compounds with an asterisk were also found in the filter blank at a lower concentration

1 **Table 1.**

Compound	CAS No	Molecular Formula	MF	RMF	tR1	tR2	Industrial Use/Source
Linear and branched OVOCs							
Propanedioic acid, propyl-	616-62-6	C ₆ H ₁₀ O ₄	885	903	18.734	0.629	Stabiliser for PVC industry / preparation of metal derivatives
Hexanoic acid, 2-ethyl-	149-57-5	C ₈ H ₁₆ O ₂	906	906	23.334	0.704	
Ethanol, 2-(2-butoxyethoxy)-	112-34-5	C ₈ H ₁₈ O ₃	920	923	26.000	1.081	Paints, cleaning, automotive products
Dipropylene glycol monomethyl ether	34590-94-8	C ₇ H ₁₆ O ₃	813	829	29.134	1.081	Paint solvents
Ethanol, 2-(2-butoxyethoxy)-, acetate	124-17-4	C ₁₀ H ₂₀ O ₄	885	886	30.000	1.333	Buildings / furnishings / paint solvents
Vanillin	121-33-5	C ₈ H ₈ O ₃	839	885	31.600	2.639	Food industry
2,4,7,9-Tetramethyl-5-decyn-4,7-diol	126-86-3	C ₁₄ H ₂₆ O ₂	892	894	31.934	0.955	Non-ionic surfactant
2,2,4-Trimethyl-1,3-pentanediol diisobutyrate	6846-50-0	C ₁₆ H ₃₀ O ₄	816	824	36.600	0.327	Food Packaging
Propanoic acid, 2-methyl-, 1-(1,1-dimethylethyl)-2-methyl-1,3-propanediyl ester	74381-40-1	C ₁₆ H ₃₀ O ₄	840	892	36.734	1.081	Plasticizer, PVC
Octane, 1,1'-oxybis-	629-82-3	C ₁₆ H ₃₄ O	904	906	37.867	0.88	Irradiated oil emissions
Isopropyl Myristate	110-27-0	C ₁₇ H ₃₄ O ₂	903	903	41.000	1.131	Cosmetic and topical medicinal preps.
2-Pentadecanone, 6,10,14-trimethyl-	502-69-2	C ₁₈ H ₃₆ O	912	915	41.534	1.181	Component of essential oil of plants
Bis(2-ethylhexyl) maleate	142-16-5	C ₂₀ H ₃₆ O ₄	786	828	46.934	1.911	Paints
Hexadecanoic acid, butyl ester	111-06-8	C ₂₀ H ₄₀ O ₂	881	881	47.534	1.634	Soaps / cosmetics / food industry / plants
2-Propenoic acid, tridecyl ester	3076-04-8	C ₁₆ H ₃₀ O ₂	857	869	49.334	1.684	Adhesives / inks
Octadecanoic acid, 2-methylpropyl ester	646-13-9	C ₂₂ H ₄₄ O ₂	841	859	50.800	1.684	Cosmetic products
Hexanedioic acid, bis(2-ethylhexyl) ester	103-23-1	C ₂₂ H ₄₂ O ₄	893	895	50.934	1.91	Plasticizer
Tetracosanoic acid, methyl ester	2442-49-1	C ₂₅ H ₅₀ O ₂	874	877	55.600	2.187	Food / beverage additives
Glycerol tricaprylate	538-23-8	C ₂₇ H ₅₀ O ₆	789	834	58.667	3.066	Pharmaceutical
Substituted aromatic OVOCs							
Ethanone, 2-(formyloxy)-1-phenyl-	55153-12-3	C ₉ H ₈ O ₃	845	960	28.334	1.935	Oil of turpentine
Terpin Hydrate	2451-01-6	C ₁₀ H ₂₂ O ₃	912	914	29.400	1.534	
Ethanone, 1,1'-(1,4-phenylene)bis-	1009-61-6	C ₁₀ H ₁₀ O ₂	873	897	32.267	2.614	Oxidation product of terpenes
cis-Pinonic acid	473-72-3	C ₁₀ H ₁₆ O ₃	884	907	32.334	1.885	
Ethanone, 1-[4-(1-hydroxy-1-methylethyl)phenyl]-	54549-72-3	C ₁₁ H ₁₄ O ₂	847	855	33.400	2.362	Rubber and Plastics
α,α,α',α'-Tetramethyl-1,4-benzenedimethanol	2948-46-1	C ₁₂ H ₁₈ O ₂	864	888	34.134	2.136	
2,6-Dimethoxybenzoquinone	530-55-2	C ₈ H ₈ O ₄	805	826	34.934	0.101	Plants / anti-cancer
Cyclopentaneacetic acid, 3-oxo-2-pentyl-, methyl ester	24851-98-7	C ₁₃ H ₂₂ O ₃	837	855	37.400	1.96	Plants / fragrance
3,5-di-tert-Butyl-4-hydroxybenzaldehyde	1620-98-0	C ₁₅ H ₂₂ O ₂	848	894	40.334	2.086	Food additive
2-Propenoic acid, 3-(4-methoxyphenyl)-, 2-ethylhexyl ester	5466-77-3	C ₁₈ H ₂₆ O ₃	867	892	47.334	2.815	Cosmetics
Phenol, 4,4'-(1-methylethylidene)bis-	80-05-7	C ₁₅ H ₁₆ O ₂	771	846	48.334	0.528	Plastics
Cyclic-OVOCs							
Succinic anhydride	108-30-5	C ₄ H ₄ O ₃	880	900	18.467	3.244	Paper production
4,8,12,16-Tetramethylheptadecan-4-olide	96168-15-9	C ₂₁ H ₄₀ O ₂	853	871	50.467	2.238	Water repellent
7,9-Di-tert-butyl-1-oxaspiro(4,5)deca-6,9-diene-2,8-dione	82304-66-3	C ₁₇ H ₂₄ O ₃	917	943	43.134	2.312	Plastic materials
1H-Naphtho[2,1-b]pyran, 3-ethenyl-dodecahydro-3,4a,7,7,10a-pentamethyl-, [3R-(3α,4αβ,6αα,10αβ,10bα)]-	596-84-9	C ₂₀ H ₃₄ O	915	918	45.800	1.985	

2

3

4

1
2
3
4
5
6
7
8
9
10
11
12
13
14
15
16
17
18
19
20
21
22
23
24
25
26
27
28
29

Table 2.

Compound	CAS No.	Molecular Formula	MF	RMF	tR1	tR2	Industrial Use/Source
OPAHs & Alicyclic Compounds							
1H-Phenalen-1-one	548-39-0	C ₁₃ H ₈ O	899	899	43.800	0.603	Dyes, paper making, pesticides Dyes, pigments
9,10-Anthracenedione	84-65-1	C ₁₄ H ₈ O ₂	898	900	45.067	0.277	
1,8-Naphthalic anhydride	81-84-5	C ₁₂ H ₆ O ₃	857	860	46.134	1.232	
9,10-Anthracenedione, methyl-	84-54-8	C ₁₅ H ₁₀ O ₂	863	877	47.467	0.327	
Naphthalene, 1-(phenylmethoxy)-	607-58-9	C ₁₇ H ₁₄ O	931	934	48.000	0.025	Heat sensitive paper
2H-Phenanthro[9,10-b]pyran		C ₁₇ H ₁₂ O	758	848	50.200	3.946	
8H-Naphtho[1,2-b]pyran-8-one, 3-ethyldecahydro-3,4a,7,7,10a-pentamethyl-, [3R-(3α,4aβ,6aa,10aβ,10bα)]-	55836-76-5	C ₂₀ H ₃₄ O ₂	754	807	50.467	3.091	
2-Ethyl-trans-4a,cis-4b,trans-8a,cis-10a-perhydro-trans-2,4a,8a-trimethylphenanthrene		C ₁₉ H ₃₂ O	780	814	50.667	3.795	
1-Phenanthrenecarboxylic acid, 1,2,3,4,4a,9,10,10a-octahydro-1,4a-dimethyl-7-(1-methylethyl)-, methyl ester, [1R-(1α,4aβ,10aa)]-	1740-19-8	C ₂₀ H ₂₈ O ₂	891	896	50.934	3.218	Pine resin
7-Oxodehydroabietic acid, methyl ester	17751-36-9	C ₂₁ H ₂₈ O ₃	835	847	54.534	0.327	Terpenic resin
16βH-Kauran-16-ol	5354-44-9	C ₂₀ H ₃₄ O	803	813	49.934	2.791	Plants

Table 3.

Compound	CAS No.	Molecular Formula	MF	RMF	tR1	tR2	Industrial Use/Source
Linear and branched ONs							
Urea, triethyl-	19006-59-8	C ₇ H ₁₆ N ₂ O	764	812	25.867	1.634	
Hexadecanamide	629-54-9	C ₁₆ H ₃₃ NO	845	845	47.400	2.438	
Substituted aromatic ONs							
Penoxaline	40487-42-1	C ₁₃ H ₁₉ N ₃ O ₄	881	885	45.867	2.966	Pesticide
Cyclic-ONs							
2,5-Pyrrolidinedione	123-56-8	C ₄ H ₅ NO ₂	864	869	22.534	2.741	Silver plating / antibacterial agent
Caprolactam	105-60-2	C ₆ H ₁₁ NO	913	913	27.267	2.967	Nylon
1,3-Diethyl-2,4,5-trioxoimidazolidine		C ₇ H ₁₀ N ₂ O ₃	881	884	28.267	2.263	Plant Metabolism
1-Ethyl-4,4-dimethyl-2,5-dioxoimidazolidine		C ₇ H ₁₂ N ₂ O ₂	808	811	28.400	2.439	
Quinoline, 2-methyl-	91-63-4	C ₁₀ H ₉ N	798	852	29.600	2.137	Dyes, Solvent for resins and terpenes

1 **Table 4.**

2

Compound	CAS No.	Molecular Formula	MF	RMF	tR1	tR2	Industrial Use/Source
Linear and branched VOCs							
Nonadecane, 2-methyl-	1560-86-7	C ₂₀ H ₄₂	820	824	40.000	0.729	Petroleum Lubricating oils
Hexadecane, 2,6,10,14-tetramethyl-	638-36-8	C ₂₀ H ₄₂	892	895	41.000	0.805	
Hexatriacontane (C ₃₆)	630-06-8	C ₃₆ H ₇₄	800	802	61.134	2.538	
Substituted aromatic VOCs							
1,1':3',1''-Terphenyl, 5'-phenyl-	612-71-5	C ₂₄ H ₁₈	887	919	60.267	3.493	
Polycyclic aromatic VOCs							
Stigmastane	601-58-1	C ₂₉ H ₅₂	791	801	62.734	0.327	

3

4

5

6

7

8

9

10

11

12

13

14

15

16

17

18

19

20

21

22

23

24

25

26

27

28

29

30

31

32

33

34

35

36

37

38

39

40

41

Table 5.

5
6
7
8
9

FIGURE LEGENDS

Figure 1. (A) Total ion current chromatogram of an urban aerosol sample (Birmingham, UK) examined by GC×GC-TOFMS illustrating improved chromatographic resolution by GC×GC compared with GC/MS (white line on background). (B) Contour plot of total ion current chromatogram of an urban aerosol sample examined by GC×GC-TOFMS. Colours indicate the intensities of the response (green>yellow>pink). (C) Example selected ion chromatogram (m/z 85, *n*-alkane fragment). Red lines signify the ordered appearance of compounds with similar physicochemical properties. Solid line – *n*-alkanes, dashed line – *n*-alkanoic acids, dotted line – furanones with varying alkyl chain length. Sample taken 09-02-11.

Figure 2. Example mass spectra of naturally occurring compounds positively identified in urban aerosol samples (Birmingham, UK). Compounds (A) and (B) are compound and reference mass spectra for 1-phenanthrenecarboxylic acid, 1,2,3,4,4a,9,10,10a-octahydro-1,4a-dimethyl-7-(1-methylethyl)-, methyl ester, [1R-(1 α ,4 $\alpha\beta$,10 $\alpha\alpha$)]; (C) and (D) are compound and reference mass spectra for 7-oxodehydroabietic acid, methyl ester; (E) and (F) are compound and reference mass spectra for cyclopentaneacetic acid, 3-oxo-2-pentyl-, methyl ester; and (G) and (H) are compound and reference mass spectra for cis-pinonic acid. (Reference spectra reproduced with permission from the National Institute of Standards & Technology spectral database ver. 2.0f Oct8 2008).

Figure 3. Example mass spectra of industrial compounds emitted from plastics, rubber and PVC, positively identified in urban aerosol samples (Birmingham, UK). Compounds (A) and (B) are compound and reference mass spectra for hexanoic acid, 2-ethyl-; (C) and (D) are compound and reference mass spectra for $\alpha,\alpha,\alpha',\alpha'$ -tetramethyl-1,4-benzenedimethanol; (E) and (F) are compound and reference mass spectra for 7,9-di-tert-butyl-1-oxaspiro(4,5)deca-6,9-diene-2,8-dione; and (G) and (H) are compound and reference mass spectra for caprolactam. (Reference spectra reproduced with permission from the National Institute of Standards & Technology spectral database ver. 2.0f Oct8 2008).

Figure 4. Example mass spectra of pesticides, flame retardants and compounds used in paper production industries, positively identified in urban aerosol samples (Birmingham, UK). Compounds (A) and (B) are compound and reference mass spectra for naphthalene, 1-(phenylmethoxy)-; (C) and (D) are compound and reference mass spectra for 2,4,7,9-Tetramethyl-5-decyn-4,7-diol; (E) and (F) are compound and reference mass spectra for penoxilane; and (G) and (H) are compound and reference mass spectra for tri(2-chloroethyl) phosphate. (Reference spectra reproduced with permission from the National Institute of Standards & Technology spectral database ver. 2.0f Oct8 2008).

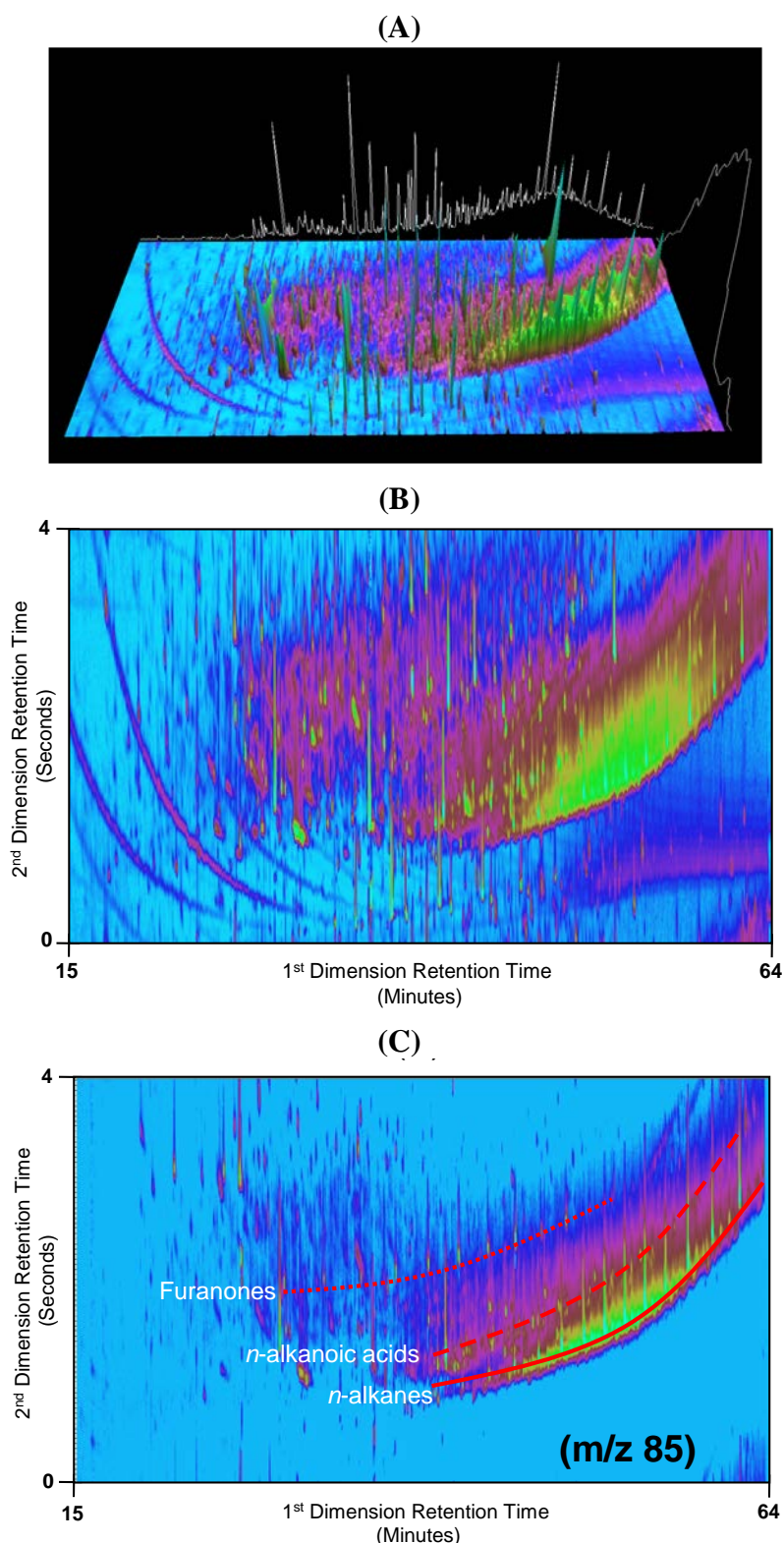


Figure 1. (A) Total ion current chromatogram of an urban aerosol sample (Birmingham, UK) examined by GC×GC-TOFMS illustrating improved chromatographic resolution by GC×GC compared with GC/MS (white line on background). (B) Contour plot of total ion current chromatogram of an urban aerosol sample examined by GC×GC-TOFMS. Colours indicate the intensities of the response (green>yellow>pink). (C) Example selected ion chromatogram (m/z 85, n -alkane fragment). Red lines signify the ordered appearance of compounds with similar physicochemical properties. Solid line – n -alkanes, dashed line – n -alkanoic acids, dotted line – furanones with varying alkyl chain length. Sample taken 09-02-11.

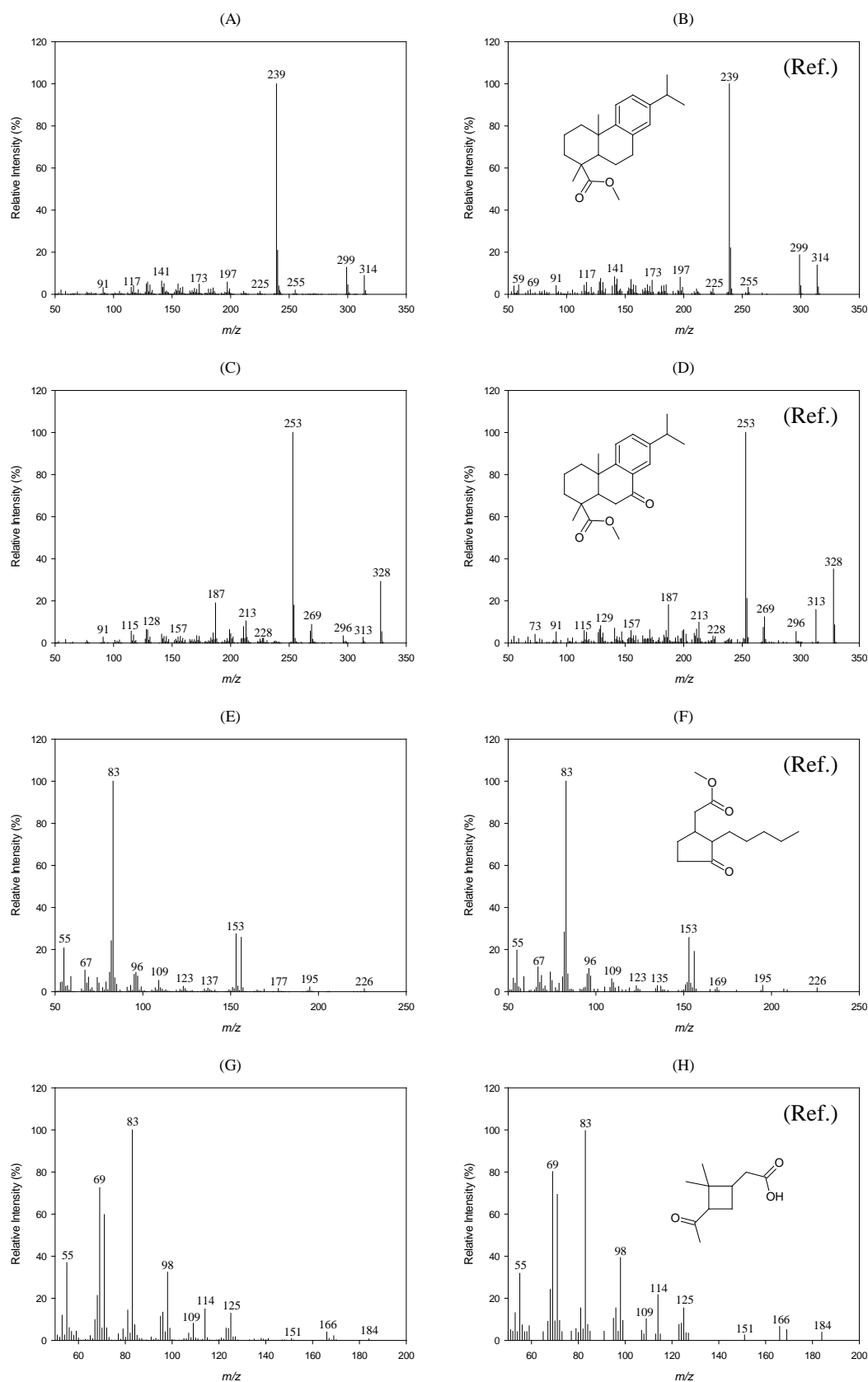


Figure 2. Example mass spectra of naturally occurring compounds positively identified in urban aerosol samples (Birmingham, UK). Compounds (A) and (B) are compound and reference mass spectra for 1-phenanthrenecarboxylic acid, 1,2,3,4,4a,9,10,10a-octahydro-1,4a-dimethyl-7-(1-methylethyl)-, methyl ester, [1R-(1 α ,4 α β ,10 α)] (Ref.); (C) and (D) are compound and reference mass spectra for 7-oxodehydroabiatic acid, methyl ester; (E) and (F) are compound and reference mass spectra for cyclopentaneacetic acid, 3-oxo-2-pentyl-, methyl ester; and (G) and (H) are compound and reference mass spectra for cis-pinonic acid. (Reference spectra reproduced with permission from the National Institute of Standards & Technology spectral database ver. 2.0f Oct8 2008).

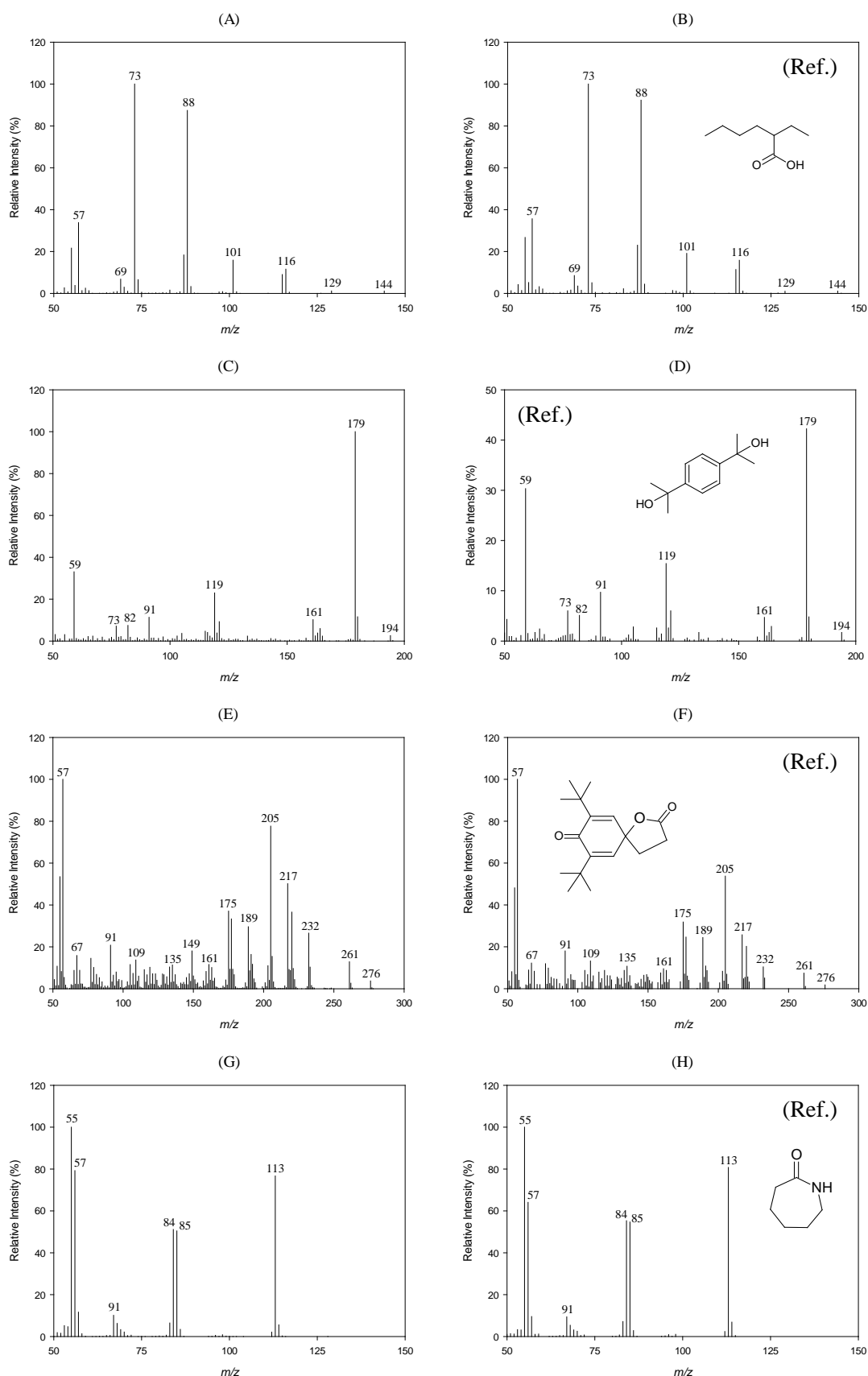


Figure 3. Example mass spectra of industrial compounds emitted from plastics, rubber and PVC, positively identified in urban aerosol samples (Birmingham, UK). Compounds (A) and (B) are compound and reference mass spectra for hexanoic acid, 2-ethyl-; (C) and (D) are compound and reference mass spectra for α, α', α' -tetramethyl-1,4-benzenedimethanol; (E) and (F) are compound and reference mass spectra for 7,9-di-tert-butyl-1-oxaspiro(4,5)deca-6,9-diene-2,8-dione; and (G) and (H) are compound and reference mass spectra for caprolactam. (Reference spectra reproduced with permission from the National Institute of Standards & Technology spectral database ver. 2.0f Oct8 2008).

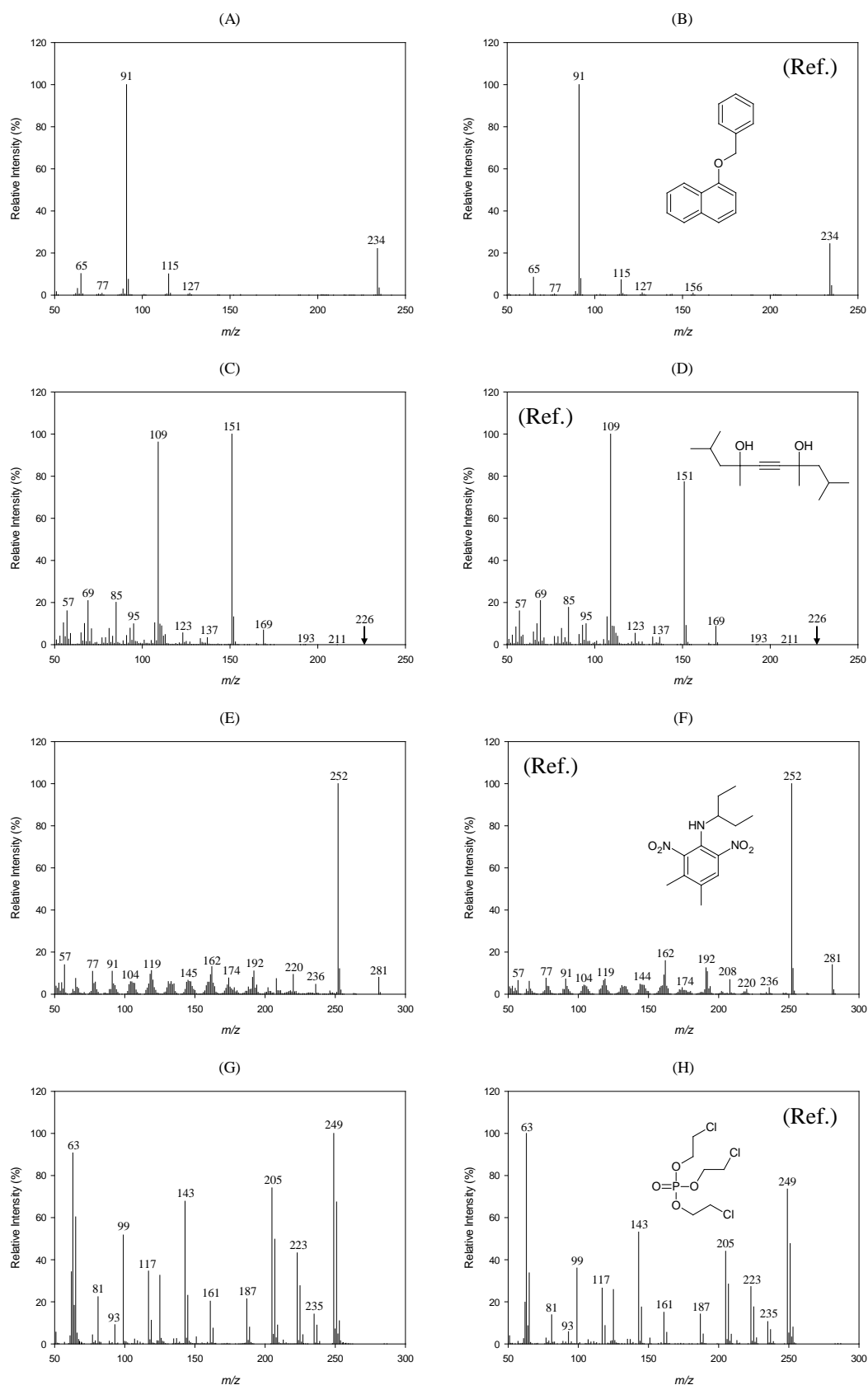


Figure 4. Example mass spectra of pesticides, flame retardants and compounds used in paper production industries, positively identified in urban aerosol samples (Birmingham, UK). Compounds (A) and (B) are compound and reference mass spectra for naphthalene, 1-(phenylmethoxy)-; (C) and (D) are compound and reference mass spectra for 2,4,7,9-Tetramethyl-5-decyn-4,7-diol; (E) and (F) are compound and reference mass spectra for penoxilane; and (G) and (H) are compound and reference mass spectra for tri(2-chloroethyl) phosphate. (Reference spectra reproduced with permission from the National Institute of Standards & Technology spectral database ver. 2.0f Oct8 2008).